



CHE1031: 'Flashcards' for chemical naming

These examples may help you use the rules for naming chemical compounds.

To make flashcards:

- grab three cards, one for each naming system; &
- write the formula on one side and the name on the other.

Ionic naming

$\text{Na}^{+1}\text{O}^{-2}$	sodium <u>oxide</u>	
FeN^{-3}	iron (III) <u>nitride</u>	[roman numerals for transition ions]
$\text{Fe}(\text{NO}_3)^{-1}$	iron (I) <u>nitrate</u>	[roman numerals for transition ions]

Molecular naming

CO	<u>mono</u> carbon <u>monoxide</u>	[no mono- for the first element]
N_3O_9	<u>trinitrogen nonoxide</u>	
P_4S_7	<u>tetraphosphorous heptasulfide</u>	

Acid naming

$\text{H}^{+1}\text{S}^{-2}$	<u>hydrosulfuric acid</u>	[hydro prefix for monoatomic anions]
$\text{H}^{+1}_2(\text{SO}_3)^{-2}$	<u>sulfurous acid</u>	[-ite → -ous]
$\text{H}^{+1}_2(\text{SO}_4)^{-2}$	<u>sulfuric acid</u>	[-ate → -ic]