



CHE1031 Chemical Naming Worksheet

Identify each as either ionic, molecular, or acid compounds:

LiF

CO₂

HCl

H(NO₃)

SF₄

CaBr₂

AlCl₃

CO

CuO

Fe₂O₃

H₃(PO₄)

PCl₃

K₂O

Pb(SO₃)

H₂(SO₄)

N₂O₄

BaCl₂

CrCl₃

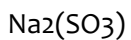
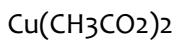
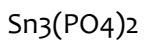
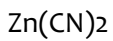
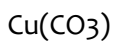
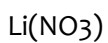
BaS

P₄O₇

BeBr₂



Name these ionic compounds:





Name these molecular compounds:

N_2O_5

PCl_5

SF_6

SO_2

PCl_3

N_2O

NO

P_4O_{10}

O_2F_2

S_4N_4

SiCl_4

ClO

ClF_3

SF_2

OF_6

SeBr_4

P_2O_5

NCl_3

P_4S_6

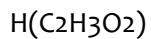
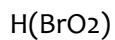
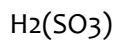
S_2Cl_2

IF_5

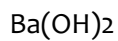
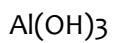
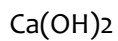
XeO_3



Name these acids:



Name these bases:





Write formulas for these mixed compounds:

calcium hydrogen sulfate

nitrous acid

potassium chloride

phosphorous pentachloride

acetic acid

iron (II) cyanide

sulfur trioxide

aluminum hydride

phosphoric acid

chromium (III) sulfate

hydrosulfuric acid

dinitrogen tetroxide

tin (II) bromide

lead (IV) sulfide

potassium chlorate

nitrogen difluoride

hydroiodic acid

nitrogen trichloride

ammonium carbonate

cesium perchlorate

silicon dioxide

ammonium hydrogen phosphate

beryllium phosphate

sulfuric acid

potassium hydrogen sulfide

calcium iodide

sulfur hexafluoride

bromous acid

silicon tetrachloride

lithium nitride

ammonium acetate

copper (II) nitrate

diphosphorous pentoxide

chlorine monoxide

sodium oxide

sodium sulfite

ruthenium (III) nitrate

potassium bromate

beryllium oxide

selenium tetrabromide

zinc (II) sulfide

calcium hydroxide

sodium selenate

lithium hydroxide