**Pre-lab quiz: Kinetics of the iodine clock reaction**

1. What is the role of starch in this lab experiment?

2. Why is Na2S2O3 included in the experiment?

3. What is the role of KCl in this experiment?

4. What is a catalyst?

5. Are these reactions being catalyzed by metal ions in this lab? Explain.

6. Describe how the concentration of rate of reaction differs in first- and second-order reactions.

6. What factors influence that rate of a chemical reaction? How?

7. In this lab experiment, what factors will be varying to change the rate of the chemical reactions?

8. Write the chemical equations involved in this experiment and show that the rate of disappearance of [S2O8-2] is proportional to the rate of appearance of the starch-iodine complex and its blue-black color.

9. How is the overall order of reaction related to the orders of reaction associated with individual reactants?