**CHE2060 Lecture 1 Quiz: Atoms, orbitals & bonding topics**

Please ask if questions are not clear to you or if you wonder whether your answer is adequate. A periodic table and table of electronegativity values is attached.

**1.1 Very brief history of the development of chemistry**

1. What are the tools that allowed for the development of modern chemistry from alchemy?

**1.2 What is organic chemistry?**

2. Define these terms:

1. Organic chemistry
2. Vitalism

**1.3 Atomic models: nuclear to quantum**

3. Valence electrons:

1. Define the term.

b. How many valence e- are there in:

N

Xe

Li

B

Ca

4. What is the major difference between the Rutherford model of the nuclear atom and the Bohr model of the quantum atom?

**1.4 All about orbitals**

5. Draw the shapes of each of these orbitals:

a. s

b. p

c. d orbital

6. In which of these orbitals does an electron possess the most potential energy?

1. 4s
2. 5s
3. 5p
4. 5d
5. 4d

7. Which orbitals have both positive and negative “lobes”?

s orbtials

p orbitals

d orbitals

**1.5 How orbitals fill: electron configuration**

8. State Hund’s Rule.

9. The term “aufbau” is a German word used to describe how electrons fill atomic orbitals by increasing energy level. At what orbital/energy level does electron filling deviate from the order you would expect? How does the actual order of filling differ from the expected order?

10. Write electron configurations for the elements listed below. Use either the abbreviated system or a box-arrow diagram.

calcium, Ca

nitrogen, N

**1.6 Basic bonding: valence electrons & molecular orbitals**

11. Two atomic orbitals combine to form:

 I. a bonding molecular orbital

 II. an antibonding molecular orbital

 III. new atomic orbtials

* 1. I only
	2. I, II and III
	3. III only
	4. I and II only

12. Why are anti-bonding orbitals rarely filled with electrons?

**1.7 Lewis dot structures of molecules**

13. Draw the Lewis structures of isocyanic acid (HNCO) and cyanic acid (HOCN).

* These structures are not resonance structures. Explain.
* When either acid loses a proton, the same ion results. Explain.

14. Complete this Lewis dot structure by adding bonds and electron pairs where needed.



**1.8 Electronegativity & bond polarity**

15. Rank these bonds in order of *increasing* polarity.

1. C – O
2. C – H
3. C – N
4. C – F
5. C – I

16. Draw the Lewis dot structures for these two molecules: CH3I and CH3F.

a. Show polarity arrows.

b. Which molecule has the larger dipole moment? Why?

c. Show the direction of the dipole moment.

17. Which of these bonds should be longest, weakest, and therefore most reactive?

C – F

C – Cl

C – Br

C – I

**1.9 Resonance: a critical concept**

18. Warfarin, the compound shown here, is marketed as the anticoagulant coumadin, and was the 18th most prescribed drug in 2006. It’s also a potent rat poison.

Circle the atom in warfarin whose formal charge that is not zero.

19. The structure of proteins depends on the amide functional group that links amino acids to form proteins. The amide group is actually a resonance structure, making for a stable protein backbone.

:O:

C

..

N

* Write both resonance structures.
* Write the resonance hybrid.

20. Which of the two resonance structures shown below makes a greater (the major) contribution to the resonance hybrid? [NB: Formal charges are not shown below!]



**1.10 Orbital hybridization: key to carbon’s “flexibility”: sp3, sp2 & sp**

21. Explain why the valence shell electrons of carbon must hybridize in order to form methane when bonding with four hydrogen atoms.

22. Draw “slot-dot” hybridization orbitals for these sp2 hybridized atoms:

sp2 carbon

sp2 nitrogen

sp2 oxygen

**1.11 Free electron pairs & radicals**

23. Consider each of these highly reactive carbon species. Calculate the formal charge on each species.

 H H H H

 | | | |

 H – C H – C . H – C : H – C :

 | | |

 H H H

**1.12 VSEPR: classifying molecular geometry & orbital hybridization**

24. Draw a detailed structure of this molecule, acetaldehyde, showing:

a. all atoms

b. all unhybridized and hybridized orbitals (including those that hold free e- pairs)

c. bond angles

 

25. Sulfur dioxide has a dipole moment, but carbon dioxide does not, even though C – O bonds are more polar than S – O bonds. Why?